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(54) Title: POLYMERIZATION OF ALKYLENE OXIDES ONTO FUNCTIONALIZED INITIATORS

(57) Abstract: Certain alcohol initiators containing halogen, aldehyde, ketone or nitro substitution can be alkoxylated with excellent efficiency and low production of by-products using a metal cyanide catalyst. The products contain halogen, aldehyde, ketone or nitro groups that can undergo subsequent reactions to form amino groups.

## POLYMERIZATION OF ALKYLENE OXIDES ONTO FUNCTIONALIZED INITIATORS

This invention relates to processes for preparing poly(oxyalkylene) polymers and to methods for preparing same.

5 Polyethers made from alkylene oxides are well known and useful in a number of applications such as detergent and cleaner compositions, oil well drilling fluids, inks, metal working fluids, lubricants in paper coating compositions, ceramics manufacturing, organic nonionic surfactants and chemical intermediates for organic nonionic surfactants which in turn are used in cosmetics, textiles and chemical processing, polyurethanes which are used  
10 as flexible foams, rigid foams and elastomers, manufacturing esters which are used in textile spin finishes, cosmetic agents, and as foam control agents for a wide variety of processes.. These polymers may have no more than one oxyalkylene group in succession, or be a higher molecular weight polymer containing one or more long chains of consecutive oxyalkylene groups.

15 Polyethers of this type are commonly made through an anionic polymerization process, whereby the alkylene oxide is combined with an initiator compound and a strongly basic catalyst such as potassium hydroxide or certain organic amines is used. The initiator compound contains one or more oxyalkylatable groups such as hydroxyl, thiol, and carboxylate. The initiator compound determines the functionality (i.e., number of hydroxyl  
20 groups/molecule of product) and in some cases may introduce some desired functional group into the product.

There are some disadvantages of polymerizing alkylene oxides using these strongly basic catalysts. Some kinds of initiator compounds cannot be alkoxylated using strongly basic catalysts because they contain base-sensitive functional groups. Examples of these  
25 are halogenated initiators and initiators that contain nitro groups. In the presence of strong bases, halogens are removed from haloalcohol compounds. Nitroalcohols react in the presence of strong bases to form a variety of by-products. In particular, nitro alcohols prepared via a Henry or Kamlet reaction or similar condensation of a nitroalkane with a carbonyl compound in the presence of base will undergo a retro reaction to generate the  
30 parent nitroalkane and carbonyl compound when exposed to basic conditions. These nitroalkane and carbonyl containing products can then undergo further base catalyzed reactions to form more undesirable by-products.

In order to try to avoid these problems, Lewis acids such as boron trifluoride-diethyl etherate and organic amines such as triethyl amine have been tried. However, some of  
35 these catalysts tend to promote the formation of large amounts of by-products, especially

when it is attempted to add three or more moles of alkylene oxide per equivalent of initiator compound. The Lewis acid catalysts tend to catalyze "back-biting" reactions where the growing polymer chain reacts with itself. The reactions form cyclic ethers such as dioxane, substituted dioxane and various crown ethers. These cannot be removed easily from the  
5 desired product, and so the product cannot be used in many applications.

Thus, it would be desirable to provide a method whereby polyethers made using halogenated or nitro-containing initiator compounds could be produced in good yield with low levels of by-products.

In one aspect, this invention is a process for preparing a polyether having one or  
10 more halogen, aldehyde, unsaturated ester, ketone or nitro groups, comprising forming a mixture of a halogenated, aldehyde-containing, unsaturated ester-containing, ketone-containing or nitro-containing initiator compound having one or one oxyalkylatable groups, at least one alkylene oxide and a metal cyanide catalyst complex, and subjecting the mixture to conditions sufficient to activate the catalyst complex and to alkoxylate the oxyalkylatable  
15 groups of the initiator.

In a second aspect, this invention is a poly(alkylene oxide) polymer containing the residue of an initiator compound containing at least one halogen, aldehyde, unsaturated ester or nitro group, the polymer having an average alkoxy degree of polymerization of at least three moles of alkylene oxide per equivalent of initiator compound.

20 This invention permits the ready formation of polymers of initiators containing halogen, aldehyde, unsaturated ester, ketone or nitro substituents that do not survive base-catalyzed alkoxylation processes. The crude polymers made by this invention do not contain significant levels of undesirable by-products such as nitroalkanes, carbonyl compounds, cyclic ethers and other impurities.

25 The polyether products of this process contain one or more terminal hydroxyl groups and at least one halogen, aldehyde, unsaturated ester, ketone or nitro group. The halogen, aldehyde, unsaturated ester, ketone or nitro functionality provides a useful reactive site on the polyether through which various types of substituents can be formed. Of principal interest is the creation of amine functionality on the polyether to form a product having both  
30 hydroxyl and amine groups.

Thus, in additional aspects this invention includes processes for making polyethers that are both amine and hydroxyl-substituted. One of these processes comprises forming a mixture of a halogenated initiator compound having one or one oxyalkylatable groups, at least one alkylene oxide and a metal cyanide catalyst complex and subjecting the mixture to  
35 conditions sufficient to activate the catalyst complex and to alkoxylate the oxyalkylatable

groups of the initiator to form a polyether containing at least one halogen group and at least one hydroxyl group, and then contacting said polyether with ammonia, a primary amine or a secondary amine under conditions sufficient to replace said halogen group with an amine group.

5 Similarly, another such process comprises forming a mixture of a nitro-containing initiator compound having one or more oxyalkylatable groups, at least one alkylene oxide and a metal cyanide catalyst complex and subjecting the mixture to conditions sufficient to activate the catalyst complex and to alkoxylation the oxyalkylatable groups of the initiator to form a polyether containing at least one nitro group and at least one hydroxyl group, and  
10 then subjecting said polyether to conditions sufficient to reduce said nitro group to an amine group.

In this invention, certain initiator compounds are alkoxylation by reaction with one or more alkylene oxides in the presence of a catalytically effective amount of a metal cyanide catalyst. The alkoxylation is conducted by combining the initiator, metal cyanide catalyst and  
15 alkylene oxide. The catalyst is then allowed to become activated in the presence of the alkylene oxide. Once the catalyst has become activated, the mixture is subjected to conditions sufficient to polymerize the alkylene oxide. In this manner, the initiator compound becomes alkoxylation until poly(oxyalkylene) chains of a desired length are introduced. As discussed below, once polymerization has begun, other types of monomers that are  
20 copolymerizable with alkylene oxides can be polymerized as well.

The initiator used herein contains at least one hydroxyl group that is bound to an aliphatic carbon atom and capable of being alkoxylation. The initiator also contains at least one halogen, aldehyde, unsaturated ester, ketone or nitro ( $-\text{NO}_2$ ) group, either of which can be bound to either an aliphatic or aromatic carbon atom. Halogen groups include fluorine,  
25 chlorine, bromine and iodine, with chlorine and bromine being preferred. The initiator may contain other functional groups as well, provided that they do not react in an undesirable way under the conditions of the alkoxylation reaction. Suitable halogenated alcohols include 2-chloroethanol, 2-bromoethanol, 2-chloro-1-propanol, 3-chloro-1-propanol, 3-bromo-1-propanol, 1,3-dichloro-2-propanol, 1-chloro-2-methyl-2-propanol, 3-chloro-2,2-dimethyl-1-propanol, 3-bromo-2,2-dimethyl-1-propanol, 4-chloro-1-butanol, 6-chloro-1-hexanol, 6-bromo-1-hexanol, 3-bromo-2-methyl-1-propanol, 7-bromo-1-heptanol, 8-chloro-1-octanol, 8-bromo-1-octanol, 2,2-dichloroethanol, 2,3-dibromopropanol, 2,2-bis(chloromethyl)-1-propanol, 2,2,2-tribromoethanol, 2,2,2-trichloroethanol, 2,2,2-trifluoroethanol, 2,2,3,3-tetrafluoro-1-propanol, 1-chloro-2-propanol, 1-bromo-2-propanol, 1,3-difluoro-2-propanol,  
30 1,3-dibromo-2-propanol, 1,4-dibromo-2-butanol, 3-chloro-1,2-propanediol, 3-bromo-1,2-

propanediol, 2-chloro-2-propene-1-ol, 2-chlorocyclohexanol, alpha-(chloromethyl)2,4-dichlorobenzyl alcohol, and trans-2,3-dibromo-2-butene-1,4-diol. Suitable ketone-containing groups include acetol and 3'-hydroxyacetophenone. Suitable nitro alcohols include 2-nitro-2-methyl-1-propanol, 2-nitroethanol, 2-nitro-1-propanol, 3-nitro-2-butanol, 3-nitro-2-pentanol, 5 2-bromo-2-nitro-1,3-propanediol, nitromethane, trispropanol, and tri(hydroxymethyl)nitromethane. Suitable unsaturated esters are compounds that are devoid of polymerizable carbon-carbon unsaturation, examples of which include ethyl glycolate and ethyl 3-hydroxybutyrate.

The alkoxylation is performed by first mixing the initiator, catalyst and an alkylene oxide and allowing the mixture to sit for a period of time at room or an elevated temperature. 10 When these materials are mixed, a so-called induction period occurs, during which the oxyalkylation reaction occurs very slowly. The induction period may range from a few minutes to several hours, depending on the particular catalyst that is used and the temperature. During this induction period, the catalyst becomes activated, and rapid 15 polymerization of the alkylene oxide then commences.

The starting mixture of catalyst, initiator compound and alkylene oxide is conveniently made by combining the catalyst and initiator compound in a pressure reactor (or by forming the catalyst in the initiator), and then pressurizing the reactor with an initial quantity of alkylene oxide. The induction period follows, as indicated by a nearly constant or slowly 20 decreasing pressure in the reactor. The onset of rapid polymerization that follows the induction period is evidenced by a drop in pressure as the alkylene oxide is consumed.

For halogenated initiators, the starting mixture of catalyst, initiator compound and alkylene oxide may be brought to any convenient temperature, such as from about 20°C, preferably from about 50°C, more preferably from about 70°C, even more preferably from 25 about 80°C to about 150°C, preferably to about 130°C. These temperatures are also suitable for conducting the polymerization once the catalyst is activated. Somewhat lower temperatures are preferred when nitro-substituted initiators are used.

Depending on the desired degree of alkoxylation, all the necessary alkylene oxide may be added to the reactor at the outset. It is usually preferred to add more alkylene oxide 30 to the reactor once the catalyst has become activated, especially when making higher molecular weight polyethers. A convenient way of adding the alkylene oxide is to pressurize the reactor with alkylene oxide and allow alkylene oxide to feed to the reactor on demand, maintaining a more or less constant pressure inside the reactor. Alternatively, any additional alkylene oxide may be fed in one or more discrete increments.

The total amount of alkylene oxide that is fed will depend on the desired equivalent weight of the product. As few as one mole of alkylene oxide per equivalent of initiator compound can be added. This invention is particularly suited for polymerizing at least about 3 moles of alkylene oxide per equivalent of initiator compound. Sufficient alkylene oxide can be added to make any desirable molecular weight polyether, such as one having a weight average molecular weight of 200,000 daltons or more. However, in most cases the intended end-use of the product will dictate its molecular or equivalent weight. For surfactant applications, molecular weights of from about 350 to about 6000 are of particular interest. In many applications, it is desirable that the product be a liquid. Poly(oxyethylene) homopolymers tend to form solids when their weight average molecular weights exceed about 700 daltons. Thus, when a poly(ethylene oxide) homopolymer is made in accordance with the invention, preferred molecular weights are about 1000 or below. All weights reported above are number average molecular weights.

Similarly, the selection of alkylene oxide will depend to a large extent on the intended end-use of the product. Among the alkylene oxides that can be polymerized with the catalyst complex of the invention are ethylene oxide, propylene oxide, 1,2-butylene oxide, styrene oxide, and mixtures thereof. Mixtures of these can be used, and two or more of them can be polymerized sequentially to make block copolymers. For polyurethanes applications, preferred alkylene oxides are propylene oxide alone, mixtures of at least 50 weight % propylene oxide and up to about 50 weight % ethylene oxide (to form a random copolymer), and propylene oxide followed by ethylene oxide, so as to form terminal poly(oxyethylene) chains constituting up to about 30% of the total weight of the product. For other applications, ethylene oxide alone, 1,2-butylene oxide, ethylene oxide/1,2-butylene oxide mixtures, ethylene oxide followed by propylene oxide or butylene oxide, butylene oxide followed by ethylene and/or propylene oxide, propylene oxide alone, mixtures of propylene oxide and ethylene and/or butylene oxide, and propylene oxide followed by ethylene and/or butylene oxide are preferred alkylene oxides.

In addition, monomers that will copolymerize with the alkylene oxide in the presence of the catalyst complex can be used to prepare modified polyether polyols, after the catalyst has become activated. Such comonomers include oxetanes as described in U. S. Patent Nos. 3,278,457 and 3,404,109 and anhydrides as described in U. S. Patent Nos. 5,145,883 and 3,538,043, which yield polyethers and polyester or polyetherester polyols, respectively. Lactones as described in U. S. Patent No. 5,525,702 and carbon dioxide are examples of other suitable monomers that can be polymerized with the catalyst of the invention.

The polymerization reaction may be performed continuously or batchwise. In such continuous processes, initiator, catalyst and alkylene oxide are continuously fed into a continuous reactor such as a continuously stirred tank reactor (CSTR) or a tubular reactor. The product continuously removed.

5 The concentration of the catalyst is selected to polymerize the alkylene oxide at a desired rate or within a desired period of time. Generally, a suitable amount of catalyst is from about 5 to about 10,000 parts by weight metal cyanide catalyst complex per million parts of the product. For determining the amount of catalyst complex to use, the weight of the product is generally considered to equal the combined weight of alkylene oxide and  
10 initiator, plus any comonomers that may be used. More preferred catalyst complex levels are from about 10, especially from about 25, to about 5000, more preferably about 1000 ppm, most preferably about 100 ppm, on the same basis.

The metal cyanide catalyst can be represented by the general formula:



wherein M is a metal ion that forms an insoluble precipitate with the  $M'(CN)_r(X)_t$  group and which has at least one water soluble salt;

$M'$  and  $M^2$  are transition metal ions that may be the same or different;

each X independently represents a group other than cyanide that coordinates with an  $M'$  or  
20  $M^2$  ion;

L represents an organic complexing agent;

$M^3A_y$  represents a water-soluble salt of metal ion  $M^3$  and anion A, wherein  $M^3$  is the same as or different than M;

b and c are positive numbers that, together with d, reflect an electrostatically neutral  
25 complex;

d is zero or a positive number;

x and y are numbers that reflect an electrostatically neutral salt;

r is from 4 to 6; t is from 0 to 2;

z is zero or a positive number and n is a positive number indicating the relative quantities of  
30 the complexing agent and  $M^3A_y$ , respectively. z and n may be fractions.

The X groups in any  $M^2(X)_e$  do not have to be all the same. The molar ratio of c:d is advantageously from about 100:0 to about 20:80, more preferably from about 100:0 to about 50:50, and even more preferably from about 100:0 to about 80:20.

Similarly, the catalyst may contain two or more types of  $M'(CN)_r(X)_t$  groups and two or  
35 more types of  $M^2(X)_e$  groups.

M and M<sup>3</sup> are preferably metal ions selected from the group consisting of Zn<sup>+2</sup>, Fe<sup>+2</sup>, Co<sup>+2</sup>, Ni<sup>+2</sup>, Mo<sup>+4</sup>, Mo<sup>+6</sup>, Al<sup>+3</sup>, V<sup>+4</sup>, V<sup>+5</sup>, Sr<sup>+2</sup>, W<sup>+4</sup>, W<sup>+6</sup>, Mn<sup>+2</sup>, Sn<sup>+2</sup>, Sn<sup>+4</sup>, Pb<sup>+2</sup>, Cu<sup>+2</sup>, La<sup>+3</sup> and Cr<sup>+3</sup>. M and M<sup>3</sup> are more preferably Zn<sup>+2</sup>, Fe<sup>+2</sup>, Co<sup>+2</sup>, Ni<sup>+2</sup>, La<sup>+3</sup> and Cr<sup>+3</sup>. M is most preferably Zn<sup>+2</sup>.

Suitable anions A include halides such as chloride and bromide, nitrate, sulfate, carbonate, cyanide, oxalate, thiocyanate, isocyanate, perchlorate, isothiocyanate, and a C<sub>1-4</sub> carboxylate. Chloride ion is especially preferred.

M<sup>1</sup> and M<sup>2</sup> are preferably Fe<sup>+3</sup>, Fe<sup>+2</sup>, Co<sup>+3</sup>, Co<sup>+2</sup>, Cr<sup>+2</sup>, Cr<sup>+3</sup>, Mn<sup>+2</sup>, Mn<sup>+3</sup>, Ir<sup>+3</sup>, Ni<sup>+2</sup>, Rh<sup>+3</sup>, Ru<sup>+2</sup>, V<sup>+4</sup> and V<sup>+5</sup>. Among the foregoing, those in the plus-three oxidation state are more preferred. Co<sup>+3</sup> and Fe<sup>+3</sup> are even more preferred and Co<sup>+3</sup> is most preferred.

Preferred groups X include anions such as halide (especially chloride), hydroxide, sulfate, C<sub>1-4</sub> carbonate, oxalate, thiocyanate, isocyanate, isothiocyanate, C<sub>1-4</sub> carboxylate and nitrite (NO<sub>2</sub><sup>-</sup>), and uncharged species such as CO, H<sub>2</sub>O and NO. Particularly preferred groups X are NO, NO<sub>2</sub><sup>-</sup> and CO.

The catalyst is usually complexed with an organic complexing agent. A great number of complexing agents are potentially useful, although catalyst activity may vary according to the selection of a particular complexing agent. Examples of such complexing agents include alcohols, aldehydes, ketones, ethers, amides, nitriles, and sulfides.

Suitable alcohols include monoalcohols and polyalcohols. Suitable monoalcohols include methanol, ethanol, n-propanol, isopropanol, n-butanol, isobutanol, t-butanol, octanol, octadecanol, 3-buten-1-ol, 3-butene-1-ol, propargyl alcohol, 2-methyl-2-propanol, 2-methyl-3-buten-2-ol, 2-methyl-3-butene-2-ol, 3-buten-1-ol, 3-butene-1-ol, and 1-t-butoxy-2-propanol. Suitable monoalcohols also include halogenated alcohols such as 2-chloroethanol, 2-bromoethanol, 2-chloro-1-propanol, 3-chloro-1-propanol, 3-bromo-1-propanol, 1,3-dichloro-2-propanol, 1-chloro-2-methyl-2-propanol as well as nitroalcohols, keto-alcohols, ester-alcohols, cyanoalcohols, and other inertly substituted alcohols.

Suitable polyalcohols include ethylene glycol, propylene glycol, glycerine, 1,1,1-trimethylol propane, 1,1,1-trimethylol ethane, 1,2,3-trihydroxybutane, pentaerythritol, xylitol, arabitol, mannitol, 2,5-dimethyl-3-hexyn-2,5-diol, 2,4,7,9-tetramethyl-5-decyne-4,7-diol, sucrose, sorbitol, alkyl glucosides such as methyl glucoside and ethyl glucoside. Low molecular weight polyether polyols, particular those having an equivalent weight of about 350 or less, more preferably about 125-250, are also useful complexing agents.

Suitable aldehydes include formaldehyde, acetaldehyde, butyraldehyde, valeric aldehyde, glyoxal, benzaldehyde, and toluic aldehyde. Suitable ketones include acetone, methyl ethyl ketone, 3-pentanone, and 2-hexanone.



Suitable ethers include cyclic ethers such as dioxane, trioxymethylene and paraformaldehyde as well as acyclic ethers such as diethyl ether, 1-ethoxy pentane, bis(betachloro ethyl) ether, methyl propyl ether, diethoxy methane, dialkyl ethers of alkylene or polyalkylene glycols (such as ethylene glycol dimethyl ether, diethylene glycol dimethyl ether, triethylene glycol dimethyl ether and octaethylene glycol dimethyl ether).

Amides such as formamide, acetamide, propionamide, butyramide and valeramide are useful complexing agents. Esters such as amyl formate, ethyl formate, hexyl formate, propyl formate, ethyl acetate, methyl acetate, and triethylene glycol diacetate can be used as well. Suitable nitriles include acetonitrile, and propionitrile. Suitable sulfides include dimethyl sulfide, diethyl sulfide, dibutyl sulfide, and diamyl sulfide.

Preferred complexing agents are t-butanol, 1-t-butoxy-2-propanol, polyether polyols having an equivalent weight of about 75-350 and dialkyl ethers of alkylene and polyalkylene glycols. Especially preferred complexing agents are t-butanol, 1-t-butoxy-2-propanol, polyether polyols having an equivalent weight of 125-250 and a dimethyl ether of mono-, di- or triethylene glycol. t-Butanol and glyme (1,2-dimethoxy ethane) are especially preferred.

A silane-functional complexing agent, as described in the copending application of Wehmeyer, application no. 09/574,842 entitled Method for Preparing Metal Cyanide Catalysts using Silane-functional Ligands, filed May 19, 2000, may be used instead of or in addition to the aforementioned complexing agents. As described therein, the silane-functional complexing agent may be polymerized to form a film or polymer, optionally on a support, or may function as a coupling agent to attach the catalyst complex to a support material.

In addition, the catalyst complex often contains a quantity of water that is bound into the crystalline lattice of the complex. Although the amount of bound water is difficult to determine, it is believed that this amount is typically from about 0.25 to about 3 moles of water per mole of  $M^1$  and  $M^2$  ions.

Exemplary catalysts include:

- Zinc hexacyanocobaltate • zL • aH<sub>2</sub>O • nZnCl<sub>2</sub>;  
 Zn[Co(CN)<sub>5</sub>NO] • zL • aH<sub>2</sub>O • nZnCl<sub>2</sub>;  
 Zn<sub>s</sub>[Co(CN)<sub>6</sub>]<sub>o</sub>[Fe(CN)<sub>5</sub>NO]<sub>p</sub> • zL • aH<sub>2</sub>O • nZnCl<sub>2</sub> (o, p = positive numbers, s=1.5o + p);  
 Zn<sub>s</sub>[Co(CN)<sub>6</sub>]<sub>o</sub>[Co(NO<sub>2</sub>)<sub>6</sub>]<sub>p</sub>[Fe(CN)<sub>5</sub>NO]<sub>q</sub> • zL • aH<sub>2</sub>O • nZnCl<sub>2</sub> (o, p, q = positive numbers, s=1.5(o+p)+q);  
 Zinc hexacyanocobaltate • zL • aH<sub>2</sub>O • nLaCl<sub>3</sub>;  
 Zn[Co(CN)<sub>5</sub>NO] • zL • aH<sub>2</sub>O • nLaCl<sub>3</sub>;  
 Zn[Co(CN)<sub>6</sub>]<sub>o</sub>[Fe(CN)<sub>5</sub>NO]<sub>p</sub> • zL • aH<sub>2</sub>O • nLaCl<sub>3</sub> (o, p = positive numbers, s=1.5o + p);

$\text{Zn}_s[\text{Co}(\text{CN})_6]_o[\text{Co}(\text{NO}_2)_6]_p[\text{Fe}(\text{CN})_5\text{NO}]_q \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{LaCl}_3$  (o, p, q = positive numbers,  $s=1.5(o+p)+q$ );

Zinc hexacyanocobaltate  $\cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{CrCl}_3$ ;

$\text{Zn}[\text{Co}(\text{CN})_5\text{NO}] \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{CrCl}_3$ ;

5  $\text{Zn}_s[\text{Co}(\text{CN})_6]_o[\text{Fe}(\text{CN})_5\text{NO}]_p \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{CrCl}_3$  (o, p = positive numbers,  $s=1.5o + p$ );

$\text{Zn}_s[\text{Co}(\text{CN})_6]_o[\text{Co}(\text{NO}_2)_6]_p[\text{Fe}(\text{CN})_5\text{NO}]_q \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{CrCl}_3$  (o, p, q = positive numbers,  $s=1.5(o+p)+q$ );

Magnesium hexacyanocobaltate  $\cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{ZnCl}_2$ ;

$\text{Mg}[\text{Co}(\text{CN})_5\text{NO}] \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{ZnCl}_2$ ;

10  $\text{Mg}_s[\text{Co}(\text{CN})_6]_o[\text{Fe}(\text{CN})_5\text{NO}]_p \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{ZnCl}_2$  (o, p = positive numbers,  $s=1.5o + p$ );

$\text{Mg}_s[\text{Co}(\text{CN})_6]_o[\text{Co}(\text{NO}_2)_6]_p[\text{Fe}(\text{CN})_5\text{NO}]_q \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{ZnCl}_2$  (o, p, q = positive numbers,  $s=1.5(o+p)+q$ );

Magnesium hexacyanocobaltate  $\cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{LaCl}_3$ ;

$\text{Mg}[\text{Co}(\text{CN})_5\text{NO}] \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{LaCl}_3$ ;

15  $\text{Mg}_s[\text{Co}(\text{CN})_6]_o[\text{Fe}(\text{CN})_5\text{NO}]_p \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{LaCl}_3$  (o, p = positive numbers,  $s=1.5o + p$ );

$\text{Mg}_s[\text{Co}(\text{CN})_6]_o[\text{Co}(\text{NO}_2)_6]_p[\text{Fe}(\text{CN})_5\text{NO}]_q \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{LaCl}_3$  (o, p, q = positive numbers,  $s=1.5(o+p)+q$ );

Magnesium hexacyanocobaltate  $\cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{CrCl}_3$ ;

$\text{Mg}[\text{Co}(\text{CN})_5\text{NO}] \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{CrCl}_3$ ;

20  $\text{Mg}_s[\text{Co}(\text{CN})_6]_o[\text{Fe}(\text{CN})_5\text{NO}]_p \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{CrCl}_3$  (o, p = positive numbers,  $s=1.5o + p$ );

$\text{Mg}_s[\text{Co}(\text{CN})_6]_o[\text{Co}(\text{NO}_2)_6]_p[\text{Fe}(\text{CN})_5\text{NO}]_q \cdot z\text{L} \cdot a\text{H}_2\text{O} \cdot n\text{CrCl}_3$  (o, p, q = positive numbers,  $s=1.5(o+p)+q$ );

as well as the various complexes such as are described at column 3 of U. S. Patent No. 3,404,109.

25 The catalyst complex may be supported. One method of making a supported catalyst is by precipitating the catalyst in the presence of a polycarboxyl or polycarboxylate compound, as described in copending application of Wehmeyer, application no. 09/574,848, entitled Method for Preparing Metal Cyanide Catalysts using Polycarboxylic Acids, filed May 19, 2000. Supported catalysts as described in WO 99/44379 are also useful. In addition,  
30 supported catalysts can be prepared as described in the copending application of Wehmeyer, application no. 09/574,843 entitled Metal Cyanide Catalysts on Inorganic Supports, filed May 19, 2000.

The catalyst complex is conveniently made using standard precipitation methods as are described, for example, in U. S. Patent Nos. 3,278,457, 3,278,458, 3,278,459,  
35 3,404,109, 3,427,256, 3,427,334, 3,427,335, 5,470,813, 5,482,908, 5,536,883, 5,589,431,

5,627,120, 5,627,122, 5,639,705, 5,714,428, 5,731,407, 5,780,584, 5,783,513. In addition, the catalyst may be formed directly as a dispersion in an initiator compound, as described in copending application of Wehmeyer, application no. 09/574,847, entitled Method for Preparing Metal Cyanide Catalyst/Polyol Initiator Slurries filed May 19, 2000, or through an incipient wetness technique as described in the copending application of Molzahn et al, U. S. Serial No. 09/574,844, entitled Incipient Wetness Method for Making Metal-Containing Cyanide Catalysts, filed May 19, 2000.

The product polyether is typically prepared in good yield with only small amounts of undesired by-products. In some instances, the product may contain a high molecular weight fraction that has a weight average molecular weight of 1.5X or more of that of the desired product. Typically, when this fraction is present, it constitutes about 20% or less, more typically less than about 10% of the total weight of the product.

Other than the high molecular weight fraction, the process of the invention permits the alkoxylation of halogenated and nitro-substituted initiator compounds with the formation of few by-products. By-products other than unreacted starting materials and the high molecular weight fraction typically constitute less than about 10% by weight of the product, more typically less than about 5 weight percent and even more typically less than about 2 weight percent, prior to any clean-up or purification. The product polyether is also generally characterized by having a good polydispersity, typically less than about 2.0, more typically less than about 1.25 and preferably less than about 1.05.

The product polyether contains at least one hydroxyl group and one or more halogen or nitro groups. The halogen and nitro groups can be reacted with various reagents to introduce various types of functionality onto the polyether. Among these reactions, amination is of particular interest. Halogen atoms on the polyether can be reacted with ammonia, a primary amine or a secondary amine to introduce the corresponding amino group ( $-NH_2$ ,  $-NHR$  or  $-NR_2$ , respectively, where R is alkyl or aryl that may be substituted). Nitro groups are easily reduced using a variety of hydrogen sources, including gaseous hydrogen over a metal catalyst, hydrochloric acid over an iron catalyst, and hydrides such as lithium aluminum hydride

The following examples are provided to illustrate the invention, but are not intended to limit its scope. All parts and percentages are by weight unless otherwise indicated.

#### Example 1

1,3-Dichloro-2-propanol (0.1121 g) and 6.2 mg of a zinc hexacyanocobaltate/t-butanol catalyst complex that has been treated with a 450 molecular weight poly(propylene

oxide) triol are charged into a dried Wheaton vial fitted with a stir bar. The vial is sealed with a septum cap and purged with nitrogen. About 0.5 g of ethylene oxide are added by syringe and the septum cap is replaced with a solid cap under nitrogen. The vial is sealed and at 90°C for 14 hours. The product has an  $M_n$  of 690 and a polydispersity of 1.22.  $C^{13}$  NMR analysis is consistent with the chloro functionality remaining intact. The peaks attributable to the carbon atoms of the initiator residue are shifted slightly from the corresponding carbon absorptions of the initiator, indicating that the 1,3-dichloro-2-propanol has initiated the polymerization.

## Example 2

2-chloroethanol (0.1161 g) and 6.3 mg of a zinc hexacyanocobaltate/t-butanol catalyst complex that has been treated with a 450 molecular weight poly(propylene oxide) triol are charged into dried Wheaton vial fitted with a stir bar. The vial is sealed with a septum cap and purged with nitrogen. About 0.5 g of ethylene oxide are added by syringe and the septum cap is replaced with a solid cap under nitrogen. The vial is sealed and at 90°C for 14 hours. The product has an  $M_n$  of 530 and a polydispersity of 1.26.  $C^{13}$  NMR analysis is consistent with the chloro functionality remaining intact. The peaks attributable to the carbon atoms of the initiator residue are shifted slightly from the corresponding carbon absorptions of the initiator, indicating that the 2-chloro-ethanol has initiated the polymerization.

## Example 3

A zinc hexacyanocobaltate/t-butanol/450 MW poly(propylene oxide) triol catalyst complex (5.94 g) and 258.84 g of 1,3-dichloro-2-propanol are homogenized and charged to a 2 gallon reactor, taking care to transfer all of the catalyst complex into the reactor. The reactor is sealed and degassed/purged several times with nitrogen, with the pressure being maintained above atmospheric pressure at all times to prevent loss of initiator. The mixture is stirred and heated to 90°C. A portion of ethylene oxide (85 g) is added. The mixture is heated and stirred. A small pressure drop is observed after about 20 minutes. Two hours after the initial addition of ethylene oxide, an additional 65 g of ethylene oxide is added. Another 515 g of ethylene oxide are then fed upon demand. The product is a yellow liquid that becomes opaque but remains liquid as it cools to room temperature. The  $M_n$  of the product is 310, and the polydispersity is 1.16. NMR analysis shows that the ethylene oxide has added onto the initiator.

**Example 4**

2-Methyl-2-nitro-1-propanol (0.1798 g) and 2 mg of a zinc hexacyanocobaltate/t-butanol catalyst complex that has been treated with a 450 molecular weight poly(propylene oxide) triol are charged into dried Wheaton vial fitted with a stir bar. The vial is sealed with a septum cap and purged with nitrogen. About 0.5 g of ethylene oxide is added by syringe and the septum cap is replaced with a solid cap under nitrogen. The vial is sealed and heated and stirred at 90°C for 14 hours. C<sup>13</sup> NMR analysis is consistent with the nitro functionality remaining intact. The peaks attributable to the carbon atoms of the initiator in the polyol are shifted from the corresponding carbon absorptions in the starting initiator, indicating that the 2-methyl-2-nitro-1-propanol has initiated the polymerization.

**Example 5**

3'-Hydroxyacetophenone (0.2005 g) and 3.7 mg of a zinc hexacyanocobaltate/t-butanol catalyst complex that has been treated with a 450 molecular weight poly(propylene oxide) triol are charged into dried Wheaton vial fitted with a stir bar. The vial is sealed with a septum cap and purged with nitrogen. About 0.5 g of ethylene oxide is added by syringe and the septum cap is replaced with a solid cap under nitrogen. The vial is sealed and heated and stirred at 90°C for 14 hours. C<sup>13</sup> NMR analysis is consistent with the ketone functionality remaining intact. The peaks attributable to the carbon atoms of the initiator in the polyol are shifted from the corresponding carbon absorptions in the starting initiator, indicating that the 3'-hydroxyacetophenone has initiated the polymerization.

**Example 6**

Acetol (0.1149 g) and 3.1 mg of a zinc hexacyanocobaltate/t-butanol catalyst complex that has been treated with a 450 molecular weight poly(propylene oxide) triol are charged into dried Wheaton vial fitted with a stir bar. The vial is sealed with a septum cap and purged with nitrogen. About 0.5 g of ethylene oxide is added by syringe and the septum cap is replaced with a solid cap under nitrogen. The vial is sealed and heated and stirred at 90°C for 14 hours. C<sup>13</sup> NMR analysis is consistent with the ketone functionality remaining intact. The peaks attributable to the carbon atoms of the initiator in the polyol are shifted from the corresponding carbon absorptions in the starting initiator, indicating that the acetol has initiated the polymerization.

**Example 7**

A monofunctional poly(ethylene oxide) is made by polymerizing 665 parts ethylene

oxide onto 258.84 parts of 1,3-dichloro-2-propanol in the presence of about 6389 ppm of a zinc hexacyanocobaltate/H<sub>2</sub>O/t-butanol catalyst complex. This corresponds to about 700 ppm Co and 1725 ppm Zn, based on the weight of the crude product to form a crude polyether. The crude product has a number average molecular weight of 310 and a polydispersity of 1.16. It contains a high molecular weight fraction constituting about 2.64% by weight of the product.

One part of the crude poly(ethylene oxide) is slurried into about two parts of isopropyl alcohol and stirred at room temperature. The liquid low molecular weight fraction mixes immediately into the solvent. About one part of n-hexane is then added, again at room temperature and the resulting mixture is stirred briefly. The mixture separates into a liquid phase and a solid phase. The solid phase is removed by vacuum filtering the mixture through filter paper and a one inch pad of a filtering aid (diatomaceous earth). The retained solids are washed with a portion of a 1:2 mixture of hexane and isopropanol. The filtered solution is then concentrated by rotary evaporation to yield the low molecular weight poly(ethylene oxide) product (about 96% recovery). The waxy solid retained on the filter bed (2.6% by weight of the total) consists of a 7400 M<sub>n</sub>, 1.22 polydispersity poly(ethylene oxide) containing substantially all of the catalyst and only a small amount of entrained low molecular weight poly(ethylene oxide).

Ten grams of the low molecular weight chloro-functionalized poly(ethylene oxide) product is combined with 70 mL of 2.0 Molar ammonia in isopropyl alcohol. The mixture is stirred and heated to 90°C in a closed vessel overnight. Substitution of the chloro groups by the amine groups is evidenced by the precipitation of ammonium chloride. C<sup>13</sup>NMR analysis is consistent with substitution of the chloro groups by amine functionality.

#### Example 8

One gram of the chloro-functionalized poly(ethylene oxide) product of Example 7 is combined with 2.0 g of 1,3-propylenediamine and stirred and heated to 90°C in a closed vessel overnight. The product is isolated by dilution with isopropyl alcohol, filtration of the precipitated hydrochloride salt and rotary evaporation of the resulting solution. C<sup>13</sup>NMR analysis is consistent with substitution of the chloro groups by amine functionality.

#### Example 9

Ethyl 3-hydroxybutyrate (0.1984 g) and 2 mg of a zinc hexacyanocobaltate/t-butanol catalyst complex that has been treated with a 450 molecular weight poly(propylene oxide) triol are charged into dried Wheaton vial fitted with a stir bar. The vial is sealed with a

septum cap and purged with nitrogen. About 0.5 g of ethylene oxide is added by syringe and the septum cap is replaced with a solid cap under nitrogen. The vial is sealed and heated and stirred at 90°C for 14 hours. C<sup>13</sup> NMR analysis is consistent with the ester functionality remaining intact. The peaks attributable to the carbon atoms of the initiator in the polyol are shifted from the corresponding carbon absorptions in the starting initiator, indicating that the alcohol hydroxy has initiated the polymerization.

#### Example 10

Ethyl glycolate (0.1561 g) and 2 mg of a zinc hexacyanocobaltate/t-butanol catalyst complex that has been treated with a 450 molecular weight poly(propylene oxide) triol are charged into dried Wheaton vial fitted with a stir bar. The vial is sealed with a septum cap and purged with nitrogen. About 0.5 g of ethylene oxide is added by syringe and the septum cap is replaced with a solid cap under nitrogen. The vial is sealed and heated and stirred at 90°C for 14 hours. C<sup>13</sup> NMR analysis is consistent with the ester functionality remaining intact. The peaks attributable to the carbon atoms of the initiator in the polyol are shifted from the corresponding carbon absorptions in the starting initiator, indicating that the alcohol hydroxy has initiated the polymerization.

1. A process for preparing a polyether having one or more halogen, aldehyde, ketone or nitro groups, comprising forming a mixture of a halogenated, aldehyde-containing, ketone-containing or nitro-containing initiator compound having one or one oxyalkylatable groups, at least one alkylene oxide and a metal cyanide catalyst complex and subjecting the mixture to conditions sufficient to activate the catalyst complex and to alkoxylate the oxyalkylatable groups of the initiator.

2. The process of claim 1 wherein the initiator compound is chlorinated or brominated.

3. The process of claim 2 wherein the alkylene oxide is ethylene oxide, propylene oxide or 1,2-butylene oxide.

4. The process of claim 3 wherein at least three moles of alkylene oxide are added per equivalent of initiator, and the polydispersity of the product is less than about 1.25.

5. The process of claim 5 wherein the initiator compound is 2-chloroethanol, 2-bromoethanol, 2-chloro-1-propanol, 3-chloro-1-propanol, 3-bromo-1-propanol, 1,3-dichloro-2-propanol, 1-chloro-2-methyl-2-propanol.

6. The process of claim 1 wherein the initiator compound contains one or more nitro groups.

7. The process of claim 6 wherein the alkylene oxide is ethylene oxide, propylene oxide or 1,2-butylene oxide.

8. The process of claim 7 wherein at least three moles of alkylene oxide are added per equivalent of initiator, and the polydispersity of the product is less than about 1.25.

9. A poly(alkylene oxide) polymer containing the residue of an initiator compound containing at least one halogen, aldehyde ketone or nitro group, the polymer having an average alkoxy degree of polymerization of at least three moles of alkylene oxide per equivalent of initiator compound.

10. The polymer of claim 9 which contains, prior to any clean-up or purification, no more than 5 weight percent of by-products other than unreacted starting materials and a



high molecular weight fraction.

11. The process of claim 9 wherein the initiator compound is chlorinated or brominated.

5        12. The process of claim 10 wherein the alkylene oxide is ethylene oxide, propylene oxide or 1,2-butylene oxide.

13. The process of claim 12 wherein at least three moles of alkylene oxide are added per equivalent of initiator, and the polydispersity of the product is less than about 1.25.

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14. The process of claim 13 wherein the initiator compound is 2-chloroethanol, 2-bromoethanol, 2-chloro-1-propanol, 3-chloro-1-propanol, 3-bromo-1-propanol, 1,3-dichloro-2-propanol, 1-chloro-2-methyl-2-propanol.

15        15. The process of claim 10 wherein the initiator compound contains one or more nitro groups.

16. The process of claim 11 wherein the alkylene oxide is ethylene oxide, propylene oxide or 1,2-butylene oxide.

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17. The process of claim 16 wherein at least three moles of alkylene oxide are added per equivalent of initiator, and the polydispersity of the product is less than about 1.25.

18. A process comprising forming a mixture of a halogenated initiator compound having one or one oxyalkylatable groups, at least one alkylene oxide and a metal cyanide catalyst complex and subjecting the mixture to conditions sufficient to activate the catalyst complex and to alkoxylate the oxyalkylatable groups of the initiator to form a polyether containing at least one halogen group and at least one hydroxyl group, and then contacting said polyether with ammonia, a primary amine or a secondary amine under conditions sufficient to replace said halogen group with an amine group.

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19. A process comprising forming a mixture of a nitro-containing initiator compound having one or one oxyalkylatable groups, at least one alkylene oxide and a metal cyanide catalyst complex and subjecting the mixture to conditions sufficient to activate the catalyst complex and to alkoxylate the oxyalkylatable groups of the initiator to form a polyether

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containing at least one nitro group and at least one hydroxyl group, and then subjecting said polyether to conditions sufficient to reduce said nitro group to an amine group.

## INTERNATIONAL SEARCH REPORT

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<b>A. CLASSIFICATION OF SUBJECT MATTER</b> IPC 7 C08G65/26 B01J27/26 C08G65/32		
According to International Patent Classification (IPC) or to both national classification and IPC		
<b>B. FIELDS SEARCHED</b> Minimum documentation searched (classification system followed by classification symbols) IPC 7 C08G B01J		
Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched		
Electronic data base consulted during the international search (name of data base and, where practical, search terms used) EPO-Internal, WPI Data, PAJ		
<b>C. DOCUMENTS CONSIDERED TO BE RELEVANT</b>		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	WO 98 27044 A (ARCO CHEM TECH NL BV ;ARCO CHEM TECH (US)) 25 June 1998 (1998-06-25) example 1 ---	1,9,10
X	EP 0 485 637 A (ASAHI GLASS CO LTD) 20 May 1992 (1992-05-20) example 11 ---	1,9,10
X	EP 0 046 947 A (CONOCO INC) 10 March 1982 (1982-03-10) claims 1,5,6 page 15 page 16 ---	9,11,12, 15,16
X	US 4 537 693 A (CAMPBELL CURTIS B) 27 August 1985 (1985-08-27) example 1 --- -/--	9,11,12, 14
<input checked="" type="checkbox"/> Further documents are listed in the continuation of box C. <input checked="" type="checkbox"/> Patent family members are listed in annex.		
* Special categories of cited documents : "A" document defining the general state of the art which is not considered to be of particular relevance "E" earlier document but published on or after the international filing date "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "O" document referring to an oral disclosure, use, exhibition or other means "P" document published prior to the international filing date but later than the priority date claimed "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art. "&" document member of the same patent family		
Date of the actual completion of the international search  18 October 2000		Date of mailing of the international search report  27/10/2000
Name and mailing address of the ISA European Patent Office, P.B. 5818 Patentlaan 2 NL - 2280 HV Rijswijk Tel. (+31-70) 340-2040, Tx. 31 651 epo nl, Fax: (+31-70) 340-3016		Authorized officer  O'Sullivan, T

# INTERNATIONAL SEARCH REPORT

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PCT/US 00/18621

## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 3 660 502 A (CASE LESLIE CATRON) 2 May 1972 (1972-05-02) example 1 ----	9, 11, 12
X	US 5 347 051 A (SHEEHAN MICHAEL T ET AL) 13 September 1994 (1994-09-13) example 2 ----	9
A		18, 19
A	EP 0 781 791 A (BASF CORP) 2 July 1997 (1997-07-02) page 5, line 39 -----	1-19

# INTERNATIONAL SEARCH REPORT

Inten... Application No  
PCT/US 00/18621

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
WO 9827044 A	25-06-1998	US 5786514 A AU 5656198 A	28-07-1998 15-07-1998
EP 0485637 A	20-05-1992	DE 69125789 D DE 69125789 T WO 9118038 A JP 4227925 A US 5290912 A	28-05-1997 14-08-1997 28-11-1991 18-08-1992 01-03-1994
EP 0046947 A	10-03-1982	JP 57045112 A	13-03-1982
US 4537693 A	27-08-1985	DE 3476679 D EP 0147438 A WO 8404754 A	16-03-1989 10-07-1985 06-12-1984
US 3660502 A	02-05-1972	NONE	
US 5347051 A	13-09-1994	US 5300691 A US 5399766 A	05-04-1994 21-03-1995
EP 0781791 A	02-07-1997	AT 168703 T CA 2193254 A DE 69600456 D DE 69600456 T ES 2119541 T CA 2193253 A CA 2193251 A US 6103850 A CA 2193252 A	15-08-1998 30-06-1997 27-08-1998 19-11-1998 01-10-1998 30-06-1997 30-06-1997 15-08-2000 30-06-1997

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For two-letter codes and other abbreviations, refer to the "Guid-  
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ning of each regular issue of the PCT Gazette.

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(54) Title: POLYMERIZATION OF ALKYLENE OXIDES ONTO FUNCTIONALIZED INITIATORS

(57) Abstract: Certain alcohol initiators containing halogen, aldehyde, ketone or nitro substitution can be alkoxyated with excellent efficiency and low production of by-products using a metal cyanide catalyst. The products contain halogen, aldehyde, ketone or nitro groups that can undergo subsequent reactions to form amino groups.

## AMENDED CLAIMS

[received by the International Bureau on 22 December 2000 (22.12.00);  
original claims 1-19 replaced by amended claims 1-17 (2 pages)]

1. A process for preparing a polyether having one or more halogen, aldehyde, ketone or nitro groups, comprising forming a mixture of a chlorinated, brominated, aldehyde-containing, ketone-containing or nitro-containing initiator compound having one or one oxyalkylatable groups, at least one alkylene oxide and a metal cyanide catalyst complex and  
5     subjecting the mixture to conditions sufficient to activate the catalyst complex and to alkoxylate the oxyalkylatable groups of the initiator.
2. The process of claim 1 wherein the alkylene oxide is ethylene oxide, propylene oxide or 1,2-butylene oxide.  
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3. The process of claim 2 wherein at least three moles of alkylene oxide are added per equivalent of initiator, and the polydispersity of the product is less than about 1.25.
4. The process of claim 1 wherein the initiator compound is 2-chloroethanol, 2-  
15     bromoethanol, 2-chloro-1-propanol, 3-chloro-1-propanol, 3-bromo-1-propanol, 1,3-dichloro-2-propanol, 1-chloro-2-methyl-2-propanol.
5. The process of claim 1 wherein the initiator compound contains one or more nitro groups.  
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6. The process of claim 5 wherein the alkylene oxide is ethylene oxide, propylene oxide or 1,2-butylene oxide.
7. The process of claim 6 wherein at least three moles of alkylene oxide are added  
25     per equivalent of initiator, and the polydispersity of the product is less than about 1.25.
8. A poly(alkylene oxide) polymer containing the residue of an initiator compound containing at least one chlorine, bromine, aldehyde, ketone or nitro group, the polymer having an average alkoxy degree of polymerization of at least three moles of alkylene oxide  
30     per equivalent of initiator compound.
9. The polymer of claim 8 which contains, prior to any clean-up or purification, no more than 5 weight percent of by-products other than unreacted starting materials and a high molecular weight fraction.

10. A polymer of claim 9 wherein the alkylene oxide is ethylene oxide, propylene oxide or 1,2-butylene oxide.

11. A polymer of claim 10 wherein at least three moles of alkylene oxide are added per equivalent of initiator, and the polydispersity of the product is less than about 1.25.

12. A polymer of claim 11 wherein the initiator compound is 2-chloroethanol, 2-bromoethanol, 2-chloro-1-propanol, 3-chloro-1-propanol, 3-bromo-1-propanol, 1,3-dichloro-2-propanol, 1-chloro-2-methyl-2-propanol.

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13. A polymer of claim 9 wherein the initiator compound contains one or more nitro groups.

14. A polymer of claim 9 wherein the alkylene oxide is ethylene oxide, propylene oxide or 1,2-butylene oxide.

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15. A polymer of claim 14 wherein at least three moles of alkylene oxide are added per equivalent of initiator, and the polydispersity of the product is less than about 1.25.

16. A process comprising forming a mixture of a halogenated initiator compound having one or one oxyalkylatable groups, at least one alkylene oxide and a metal cyanide catalyst complex and subjecting the mixture to conditions sufficient to activate the catalyst complex and to alkoxylate the oxyalkylatable groups of the initiator to form a polyether containing at least one halogen group and at least one hydroxyl group, and then contacting said polyether with ammonia, a primary amine or a secondary amine under conditions sufficient to replace said halogen group with an amine group.

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17. A process comprising forming a mixture of a nitro-containing initiator compound having one or one oxyalkylatable groups, at least one alkylene oxide and a metal cyanide catalyst complex and subjecting the mixture to conditions sufficient to activate the catalyst complex and to alkoxylate the oxyalkylatable groups of the initiator to form a polyether containing at least one nitro group and at least one hydroxyl group, and then subjecting said polyether to conditions sufficient to reduce said nitro group to an amine group.

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